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Isotopes and average atomic mass worksheet answer

Name: _____
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Atomic Mass and Atomic Number Worksheet

Name of Element	Symbol	Atomic Number	Atomic Mass	Protons	Neutrons	Electrons
Copper				29	35	29
Tin	Sn				69	50
Iodine	I	53	127			
Uranium			238			92
Potassium	K			19	20	
Lithium			7	3		
Oxygen	O	8			8	
Gold		79	197			
		16	32	16		
Silver		47	108	47		
Chromium					28	24
	Co		59		32	27
	Ni			28		
Zinc		30			35	
	Al				14	13
Mercury	Hg	80	201			
Platinum			195			
	Fe		56		30	
	H	1	1			
	He	2	4			
		4		4		4
	Mg			12	12	12
	C	6		6	6	
Silicon		14			14	
	Cl			17	18	
	Bi		209			83
Boron		5	11			
	Ca	20				20
		25	55	25		
Lead			207			82
Sodium	Na					
Fluorine				9	10	9
	P	15	31			

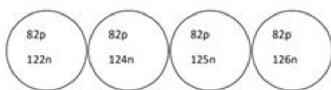
Unit 3—Atomic Theory and the Periodic Table

Name _____ Class Period _____

Pre-AP Chemistry: Worksheet #3.3

Isotopes and Average Atomic Mass

- Name two ways that isotopes of an element differ.
Mass Number, Atomic Mass, Neutrons
- What data must you know about the isotopes of an element to calculate the atomic mass of the element?
Atomic Mass of each isotope and % abundance of each isotope
- The four isotopes of lead are shown below, each with its percent by mass abundance and the composition of its nucleus. Using these data, calculate the approximate atomic mass of lead.



- Lithium has two naturally occurring isotopes. Lithium-6 has an atomic mass of 6.015 amu; lithium-7 has an atomic mass of 7.016 amu. The atomic mass of lithium is 6.941 amu. What is the percentage of naturally occurring lithium-7? (Make Li-6's percent abundance x and Li-7's percent abundance $1-x$)

Isotope	Atomic Mass	% Abundance	
Lithium-6	6.015 amu	$x = 0.0749$	$\rightarrow 7.49\%$
Lithium-7	7.016 amu	$1-x = 1 - 0.0749 = 0.9251$	$\rightarrow 92.51\%$

$6.015x + 7.016(1-x) = 6.941 \rightarrow x = 0.7492$

- What is the average atomic mass of an element, and how does it differ from the mass number?
The atomic mass of an element is the average of all of the masses of the naturally occurring isotopes of that element. The mass number represents the protons plus the neutrons for a certain isotope of that element.
- Imagine you are standing on top of a boron-11 nucleus. Describe the numbers and kinds of subatomic particles you would see looking down into the nucleus, and those you would see looking out from the nucleus.
Inside the nucleus—5 protons and 6 neutrons Outside the nucleus—6 electrons

Model 2 - Natural Abundance Information for Magnesium

Isotope	Percent Abundance (in Part 1)	Atomic Mass (amu)
^{24}Mg	78.99	23.98504
^{25}Mg	10.00	24.98584
^{26}Mg	11.01	25.98259

- Calculate the average atomic mass of magnesium from Model 2.
- Calculate the percent number of protons of each isotope and the total number of protons in a sample of 10 atoms of Mg. Why do you think the number of protons is a whole number?
- Calculate the percent number of neutrons of each isotope and the total number of neutrons in a sample of 10 atoms of Mg. Why do you think the number of neutrons is a whole number?
- Calculate the percent number of electrons of each isotope and the total number of electrons in a sample of 10 atoms of Mg. Why do you think the number of electrons is a whole number?
- Calculate the average atomic mass of magnesium from the percent number of protons and neutrons in a sample of 10 atoms of Mg. Why do you think the average atomic mass is a decimal number?
- Calculate the average atomic mass of magnesium from the percent number of protons and neutrons in a sample of 10 atoms of Mg. Why do you think the average atomic mass is a decimal number?

element and assigns an average value to the mass of one atom of the element. For example, the atomic mass of carbon is 12.011 atomic mass units generally contain 98.89% of the carbon-12 isotope, 1.11% of carbon-13, and trace amounts of carbon-14. A force acting upon an object for some duration of time results in an impulse. The quantity impulse is calculated by multiplying force and time. Impulses cause objects to change their momentum. And finally, the impulse an object experiences is equal to ... 06/10/2021 · Chemistry is a branch of science that deals with the study of matter and its chemical properties. Learn about the definition of chemistry, its rich history, and the different branches of chemistry. The LibreTexts libraries are Powered by MindTouch © and are supported by the Department of Education Open Textbook Pilot Project, the UC Davis Office of the Provost, the UC Davis Library, the California State University Affordable Learning Solutions Program, and Merlot. We also acknowledge previous National Science Foundation support under grant numbers 1246120, ... What was the average air resistance on the skydiver during this jump? How does the magnitude of the average air resistance compare to the weight of the skydiver? Top pole vaulters have a mass of about 80 kg and can clear a bar 6.0 m above the ground. Top sprinters also have a mass of about 80 kg and can cover 100 m in 10 s. Although mass has no effect on the acceleration due to gravity, there are three factors that do. They are location, location, location. Everyone reading this should be familiar with the images of the astronauts hopping about on the moon and should know that the gravity there is weaker than it is on the Earth — about one sixth as strong or 1.6 m/s 2 . The most critical question in deciding how an object will move is to ask are the individual forces that act upon balanced or unbalanced? The manner in which objects will move is determined by the answer to this question. Unbalanced forces will cause objects to change their state of motion and a balance of forces will result in objects continuing in their current state of motion. 28/02/2020 · The atomic weight values may be different from one periodic table to another because it's a calculated number, based on the weighted average of the natural isotopes of an element. If a new supply of an element is discovered, the isotope ratio may be different from what scientists previously believed. For these sorts of applications, the percent composition of a compound is easily derived from its formula mass and the atomic masses of its constituent elements. A molecule of NH 3 contains one N atom weighing 14.01 amu and three H atoms ... Although mass has no effect on the acceleration due to gravity, there are three factors that do. They are location, location, location. Everyone reading this should be familiar with the images of the astronauts hopping about on the moon and should know that the gravity there is weaker than it is on the Earth — about one sixth as strong or 1.6 m/s 2 . 2 dias atrás · Can a cop pull you over without his blue lights on answer bing, teachers guide glk123. 10. 20 Questions Show answers. QUESTION. Average score for this quiz is 9 / 15. Brainpop ethics quiz answers - mmlml. Everfi Module 3 Answers Key - Joomlaxe. Your timeline should include all 5 underlined vocabulary words from Section 3 (pgs.

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